

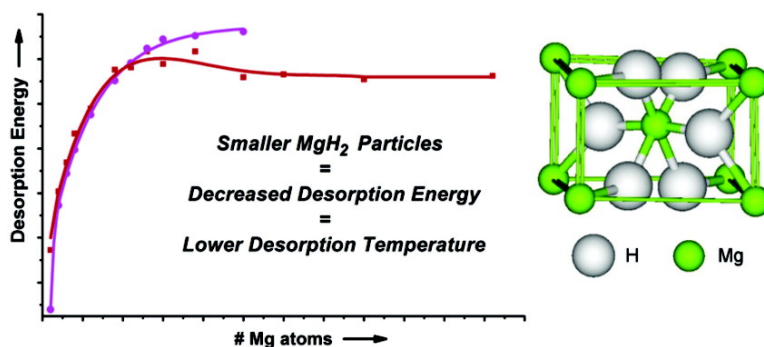
Article

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Hydrogen Storage in Magnesium Clusters: Quantum Chemical Study

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Abstract: Magnesium hydride is cheap and contains 7.7 wt % hydrogen, making it one of the most attractive hydrogen storage materials. However, thermodynamics dictate that hydrogen desorption from bulk magnesium hydride only takes place at or above 300 °C, which is a major impediment for practical application. A few results in the literature, related to disordered materials and very thin layers, indicate that lower desorption temperatures are possible. We systematically investigated the effect of crystal grain size on the thermodynamic stability of magnesium and magnesium hydride, using ab initio Hartree–Fock and density functional theory calculations. Also, the stepwise desorption of hydrogen was followed in detail. As expected, both magnesium and magnesium hydride become less stable with decreasing cluster size, notably for clusters smaller than 20 magnesium atoms. However, magnesium hydride destabilizes more strongly than magnesium. As a result, the hydrogen desorption energy decreases significantly when the crystal grain size becomes smaller than ~1.3 nm. For instance, an MgH₂ crystallite size of 0.9 nm corresponds to a desorption temperature of only 200 °C. This predicted decrease of the hydrogen desorption temperature is an important step toward the application of Mg as a hydrogen storage material.

1. Introduction

Hydrogen is an ideal clean carrier for storage, transport, and conversion of energy. However, a key problem is its storage, especially for its use as a fuel for zero-emission mobile applications. Physical storage of hydrogen, as a highly pressurized gas or in a liquid phase at low temperatures, is associated with significant security risks, energy losses, and high volumes.^{1,2} Furthermore, physical adsorption of hydrogen onto lightweight nanomaterials, such as zeolites, carbon nanotubes, and activated carbons, yields only low storage densities and also needs low temperatures.^{3–8} The alternative is chemical storage: the reversible absorption of hydrogen into another material. Especially metal hydrides offer a promising and safe alternative to storage in compressed or liquid form.^{9,10}

Hydrogen storage in metal hydrides has been the focus of intensive research. Magnesium dihydride combines a high H₂ capacity of 7.7 wt % with the benefit of the low cost of the abundantly available magnesium.^{11–13} The main barriers for direct usage of pure MgH₂ are slow desorption kinetics, a high thermodynamic stability, and a high reactivity toward air and oxygen which it has in common with most other lightweight metal hydrides.^{14,15} The high thermodynamic stability of MgH₂ results in a relatively high desorption enthalpy, which corresponds to an unfavorable desorption temperature of 573 K at 1 bar H₂.^{9,11,16} During the past two decades research efforts were devoted to modifying the Mg-based system, aiming at increasing the absorption/desorption rates and lowering the desorption temperature. Different approaches were reported, mainly involving alloying Mg with other elements, high energy or reactive ball milling of Mg, or surface modification of Mg.^{17–25} Many

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- (1) Schlapbach, L.; Züttel, A. *Nature* **2001**, *414* (6861), 353–358.
- (2) Züttel, A. *Naturwissenschaften* **2004**, *91* (4), 157–172.
- (3) Zecchina, A.; Bordiga, S.; Vitillo, J. G.; Ricchiardi, G.; Lamberti, C.; Spoto, G.; Bjorgen, M.; Lillerud, K. P. *J. Am. Chem. Soc.* **2005**, *127* (17), 6361–6366.
- (4) Hirscher, M.; Becher, M.; Haluska, M.; Dettlaff-Weglikowska, U.; Quintel, A.; Duesberg, G. S.; Choi, Y. M.; Downes, P.; Hulman, M.; Roth, S.; Stepanek, I.; Bernier, P. *Appl. Phys. A: Mater. Sci. Process.* **2001**, *72* (2), 129–132.
- (5) Terres, E.; Panella, B.; Hayashi, T.; Kim, Y. A.; Endo, M.; Dominguez, J. M.; Hirscher, M.; Terrones, H.; Terrones, M. *Chem. Phys. Lett.* **2005**, *403* (4–6), 363–366.
- (6) Dillon, A. C.; Heben, M. J. *Appl. Phys. A: Mater. Sci. Process.* **2001**, *72* (2), 133–142.
- (7) Nijkamp, M. G.; Raaymakers, J.; van Dillen, A. J.; de Jong, K. P. *Appl. Phys. A: Mater. Sci. Process.* **2001**, *72* (5), 619–623.
- (8) Schimmel, H. G.; Kearley, G. J.; Nijkamp, M. G.; Visserl, C. T.; de Jong, K. P.; Mulder, F. M. *Chem.—Eur. J.* **2003**, *9* (19), 4764–4770.

- (9) Huot, J.; Liang, G.; Schulz, R. *Appl. Phys. A: Mater. Sci. Process.* **2001**, *72* (2), 187–195.
- (10) Zaluska, A.; Zaluski, L.; Strom-Olsen, J. O. *Appl. Phys. A: Mater. Sci. Process.* **2001**, *72* (2), 157–165.
- (11) Imamura, H.; Masanari, K.; Kusuhara, M.; Katsumoto, H.; Sumi, T.; Sakata, Y. *J. Alloys Compd.* **2005**, *386* (1–2), 211–216.
- (12) Zaluski, L.; Zaluska, A.; Strom-Olsen, J. O. *J. Alloys Compd.* **1997**, *253*, 70–79.
- (13) Liang, G.; Huot, J.; Boily, S.; Van Neste, A.; Schulz, R. *J. Alloys Compd.* **1999**, *291* (1–2), 295–299.
- (14) Zaluska, A.; Zaluski, L.; Strom-Olsen, J. O. *J. Alloys Compd.* **1999**, *288* (1–2), 217–225.
- (15) Barkhordarian, G.; Klassen, T.; Bormann, R. *J. Alloys Compd.* **2004**, *364* (1–2), 242–246.
- (16) Grochala, W.; Edwards, P. P. *Chem. Rev.* **2004**, *104* (3), 1283–1315.
- (17) Vegge, T.; Hedegaard-Jensen, L. S.; Bonde, J.; Munter, T. R.; Norskov, J. K. *J. Alloys Compd.* **2005**, *386* (1–2), 1–7.

of these techniques are suitable to improve the kinetics drastically, increasing sorption rates by up to 2 orders of magnitude. This can be ascribed mainly to surface enlargement, dissolution of other metals into Mg (disruption of the Mg crystalline structure), decreasing diffusion lengths, and breaking up the inhibiting and passivating outer oxide layer. However, upscaling of milling is not straightforward and this technique is limited to grain sizes down to 10–50 nm for pure Mg. More importantly, the thermodynamics are not affected by such techniques, so the desorption temperature of MgH₂ cannot be lowered below that of the bulk value of 573 K. Although some alloying or doping techniques are able to affect the desorption temperature, this is accompanied by a lower hydrogen storage capacity due to the added weight.^{10,26–28} Furthermore, the positive effects on sorption temperatures and rates of these milled and/or doped Mg compounds are often lost in the first few cycles of charging and discharging with hydrogen. A similar effect is observed with other bulk phases of magnesium hydride.^{9,13,29–33} γ -MgH₂ is a less stable phase than the more common β -MgH₂, but the improvement in desorption temperature is lost after the first hydrogenation/dehydrogenation cycle, upon which β -MgH₂ is formed. Lower desorption temperatures have also been reported for sputtered thin films of Mg, mostly capped with a Pd layer, or for highly amorphous Mg.^{26,34} For these materials the desorption temperatures are more stable upon cycling, but this effect is not yet understood, although it might be related to the presence of the Pd.

These promising results have stimulated us to consider in more detail the thermodynamic stability of MgH₂ versus Mg + H₂ as a function of crystal grain size. The classical Born–Haber cycle for the magnesium–hydrogen system is shown in Figure 1. Although MgH₂ is not an entirely ionic structure as assumed in this thermodynamic cycle, the general scheme can be used as a first approximation.^{35–37} The lattice energy for MgH₂ is considerably larger than that for Mg and is also ca. 36 times

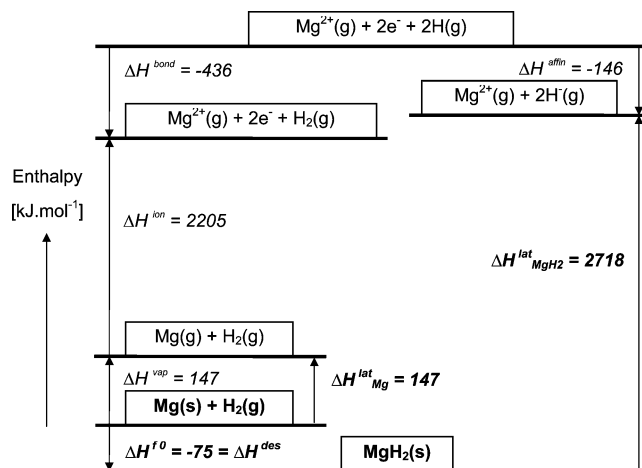


Figure 1. Born–Haber cycle for the bulk Mg–H₂ system at standard pressure and temperature (1 × 10⁵ Pa, 298 K). The lattice enthalpies [kJ·mol⁻¹] are given in bold, as is the formation enthalpy of MgH₂, which corresponds to the desorption enthalpy.

larger than the desorption enthalpy. Therefore, a relatively small shift in the lattice enthalpies can have a large impact on the desorption enthalpy. In general, the specific lattice enthalpy decreases upon lowering the amount of atoms in a cluster (Mg_x or Mg_xH_{2x}), due to the decreased average coordination number of the atoms. Distortions of the Mg(H₂) lattice can be expected to result in a shift in the lattice enthalpy, similar to what is seen for alloyed magnesium.^{17,38–41}

For main group metals, the electronic structure varies with the number of interacting atoms.^{42–45} The equilibrium geometries and electronic structures of small magnesium clusters have been described in the literature for different quantum chemical methods.^{46–54} For thin magnesium layers a decrease in potential energy was reported upon decreasing the layer thickness in one dimension.³⁹ Although quantum chemical calculations have been performed for other metal hydrides and molecules, a systematic study on the geometry and stability of MgH₂ as a function of cluster size has not yet been reported.⁵⁵ In this paper we present

(18) Au, M. *Mater. Sci. Eng., B: Solid State Mater. Adv. Technol.* **2005**, *117* (1), 37–44.
 (19) Liang, G. J. *Alloys Compd.* **2004**, *370* (1–2), 123–128.
 (20) Oelerich, W.; Klassen, T.; Bormann, R. *Mater. Trans.* **2001**, *42* (8), 1588–1592.
 (21) Huot, J.; Liang, G.; Schulz, R. *J. Alloys Compd.* **2003**, *353* (1–2), L12–L15.
 (22) Oelerich, W.; Klassen, T.; Bormann, R. *J. Alloys Compd.* **2001**, *315* (1–2), 237–242.
 (23) Huot, J.; Swainson, I. P.; Schulz, R. *J. Alloys Compd.* **1999**, *292* (1–2), 292–295.
 (24) Wang, X. L.; Suda, S. *J. Alloys Compd.* **1995**, *231* (1–2), 380–386.
 (25) Liang, G.; Huot, J.; Boily, S.; Schulz, R. *J. Alloys Compd.* **2000**, *305* (1–2), 239–245.
 (26) Higuchi, K.; Kajioka, H.; Toiyama, K.; Fujii, H.; Orimo, S.; Kikuchi, Y. *J. Alloys Compd.* **1999**, *295*, 484–489.
 (27) Chen, J.; Yang, H. B.; Xia, Y. Y.; Kuriyama, N.; Xu, Q.; Sakai, T. *Chem. Mater.* **2002**, *14* (7), 2834–2836.
 (28) Orimo, S.; Fujii, H. *Appl. Phys. A: Mater. Sci. Process.* **2001**, *72* (2), 167–186.
 (29) Schimmel, H. G. Thesis, Delft University of Technology, 2004.
 (30) Vajeeston, P.; Ravindran, P.; Kjekshus, A.; Fjellvag, H. *Phys. Rev. Lett.* **2002**, *89* (17), 1755061–1755064.
 (31) Yavari, A. R.; de Castro, J. F. R.; Vaughan, G.; Heunen, G. *J. Alloys Compd.* **2003**, *353* (1–2), 246–251.
 (32) Schimmel, H. G.; Johnson, M. R.; Kearley, G. J.; Ramirez-Cuesta, A. J.; Huot, J.; Mulder, F. M. *Mater. Sci. Eng., B: Solid State Mater. Adv. Technol.* **2004**, *108* (1–2), 38–41.
 (33) Wang, P.; Zhang, H. F.; Ding, B. Z.; Hu, Z. Q. *J. Alloys Compd.* **2000**, *313*, 209–213.
 (34) Higuchi, K.; Yamamoto, K.; Kajioka, H.; Toiyama, K.; Honda, M.; Orimo, S.; Fujii, H. *J. Alloys Compd.* **2002**, *330*, 526–530.
 (35) Ohba, N.; Miwa, K.; Noritake, T.; Fukumoto, A. *Phys. Rev. B* **2004**, *70* (3), 0351021–0351027.
 (36) Noritake, T.; Towata, S.; Aoki, M.; Seno, Y.; Hirose, Y.; Nishibori, E.; Takata, M.; Sakata, M. *J. Alloys Compd.* **2003**, *356*, 84–86.
 (37) Noritake, T.; Aoki, M.; Towata, S.; Seno, Y.; Hirose, Y.; Nishibori, E.; Takata, M.; Sakata, M. *Appl. Phys. Lett.* **2002**, *81* (11), 2008–2010.

(38) Liang, J. J.; Khosrovani, N. *Prepr. Pap.—Am. Chem. Soc., Div. Fuel Chem.* **2004**, *49* (1), 152–154.
 (39) Liang, J. J. *Appl. Phys. A: Mater. Sci. Process.* **2005**, *80* (1), 173–178.
 (40) Bach, F. W.; Schaper, M.; Jaschik, C. Influence of lithium on hcp magnesium alloys. In *Magnesium Alloys 2003, Parts 1 And 2 Materials Science Forum 419–4*; Trans Tech Publications Ltd.: Zurich, Switzerland, 2003; pp 1037–1042.
 (41) Notten, P. H. L.; Outwerkerk, M.; van Hal, H.; Beelen, D.; Keur, W.; Zhou, J.; Feil, H. *J. Power Sources* **2004**, *129* (1), 45–54.
 (42) Bandow, S.; Kumura, K. *Solid State Commun.* **1990**, *73* (2), 167–171.
 (43) Sichel, E. K.; Gittleman, J. I.; Abeles, B. *Thin Solid Films* **1978**, *51*, 89–92.
 (44) Millet, J.-L.; Borel, J.-P. *Solid State Commun.* **1982**, *43* (3), 217–220.
 (45) Belanzoni, P.; van Lenthe, E.; Baerends, E. J. *J. Chem. Phys.* **2001**, *114* (10), 4421–4433.
 (46) Diederich, T.; Doppner, T.; Braune, J.; Tiggesbaumker, J.; Meiwes-Broer, K. H. *Phys. Rev. Lett.* **2001**, *86* (21), 4807–4810.
 (47) Tsuda, M.; Dino, W. A.; Nakanishi, H.; Kasai, H. *J. Phys. Soc. Jpn.* **2004**, *73* (10), 2628–2630.
 (48) Lyalin, A.; Solovyov, I. A.; Solovyov, A. V.; Greiner, W. *Phys. Rev. A* **2003**, *67* (6), 851–859.
 (49) Thomas, O. C.; Zheng, W. J.; Xu, S. J.; Bowen, K. H. *Phys. Rev. Lett.* **2002**, *89* (21), 213403-1.
 (50) Davidson, E. R.; Frey, R. F. *J. Chem. Phys.* **1997**, *106* (6), 2331–2341.
 (51) Wang, Y. J.; Yuan, H. T.; Wang, L. B.; Cao, J. S.; Li, Q. D. *Rare Met. Mater. Eng.* **2001**, *30*, 267–270.
 (52) Griessen, R.; Driessen, A. *Phys. Rev. B* **1984**, *30* (8), 4372–4382.
 (53) Bogdanovic, B.; Bohmhammel, K.; Christ, B.; Reiser, A.; Schlichte, K.; Vehlen, R.; Wolf, U. *J. Alloys Compd.* **1999**, *282* (1–2), 84–92.
 (54) Bohmhammel, K.; Wolf, U.; Wolf, G.; Konigsberger, E. *Thermochim. Acta* **1999**, *337* (1–2), 195–199.
 (55) Smithson, H.; Marianetti, C. A.; Morgan, D.; Van der Ven, A.; Predith, A.; Ceder, G. *Phys. Rev. B* **2002**, *66* (14), 144107.

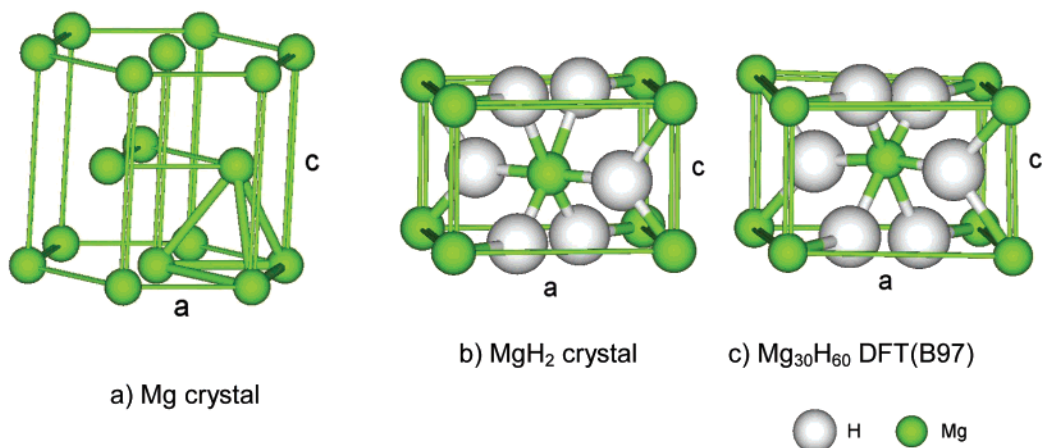
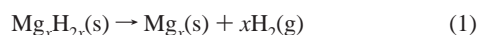


Figure 2. Structures of (a) unit cell of bulk Mg, (b) unit cell of bulk β -MgH₂, and (c) core of Mg₃₀H₆₀.

a theoretical investigation of the cluster size dependency of the (de)sorption enthalpy by calculations on Mg and MgH₂ clusters. Ab initio calculations have been applied for clusters with up to 30 magnesium atoms. For larger clusters up to 56 magnesium atoms, pseudopotentials and density functional theory methods were used. When the potential energies for both Mg and MgH₂ clusters are calculated in a consistent manner, a direct relation between cluster size and the energy can be made. By comparing the stabilities of small metallic and hydride clusters, we can predict changes in (equilibrium) thermodynamics, which can have a large impact on the desorption temperature.

2. Methods

To evaluate the desorption enthalpy, the energies of both Mg and MgH₂ clusters are needed. For all clusters considered, the geometry with the lowest possible energy was determined in a gradient-driven geometry optimization, after which the zero point vibrational energy (ZPE) was calculated for that geometry using analytical Hessians.⁵⁶ Reaction 1 represents the desorption of hydrogen from a magnesium hydride cluster:



The energies (E_{min} and ZPE) of all components involved in this reaction were calculated for a range of cluster sizes up to 56 Mg atoms. Where possible, the values for Mg were compared with those from the literature. Contrary to the solid clusters, the hydrogen molecules are in the gas phase. Since the hydrogen molecules are relatively far apart, their intermolecular interactions are negligible. The desorption enthalpy can be approximated by the desorption energy ($\Delta H^{\text{des}} \approx \Delta E^{\text{des}}$), because both the contributing $\Delta[p \, dV]$ and $\Delta[c_p \, dT]$ terms are small in comparison to the energy term over the whole reaction. The energy of the reaction was taken as the difference in calculated energies, as shown in eq 2.

$$\Delta E^{\text{des}} \approx E[\text{Mg}_x] + xE[\text{H}_2] - E[\text{Mg}_x\text{H}_{2x}] \quad (2)$$

With the ab initio restricted Hartree–Fock (HF) method, the desired energies can be calculated for clusters up to 30 Mg atoms. The used basis set was SV 6-31G,⁵⁷ and the geometries were optimized with a convergence threshold of 0.001 $E_{\text{h}} \cdot a_0^{-1}$ in the gradient.

To validate the calculation method, we verified whether the calculated desorption energies were in good agreement with the bulk experimental value of 75 kJ·mol⁻¹ [H₂]⁻¹ for large cluster sizes. With the use of a pseudopotential for Mg, the energies for clusters up to 56 Mg atoms could be calculated. For these calculations we incorporated

density functional theory (DFT) to obtain a comparison with results taken from calculations on magnesium metal clusters and the bulk values for the desorption enthalpies. For the DFT calculations, the B97 functional⁵⁸ was used with the ECP basis set with a double- ζ contraction (ECPDZ) for magnesium, and a convergence threshold of 0.002 $E_{\text{h}} \cdot a_0^{-1}$ in the gradient. The B97 functional has been shown in the past to yield reasonable results for small molecules or clusters, both with and without metal atoms. All calculations were performed with the GAMESS-UK program, version 6.0, on the TERAS computer of SARA.⁵⁹

3. Results and Discussion

3.1. Geometric Optimization and Validation. To validate the geometric optimization for both calculation methods, all structures were checked for plausibility, meaning no emerging hydrogen molecules and a narrow distribution for the Mg–Mg, H–H, and Mg–H bond lengths. Furthermore, different starting geometries were used for each cluster. The calculated geometries and energies for the Mg clusters were similar to those stated in the literature.^{39,48,50,60} The transition from the magnesium metal toward the magnesium hydride is accompanied by a large change in structure for the Mg atoms, as illustrated in Figure 2. Upon charging with hydrogen, the metallic Mg atoms trade their hexagonal environment of the hcp structure for a bcc sublattice in the rutile structure. The atomic radius of magnesium shrinks upon becoming partly cationic. This yields enough space for the large hydrogen atoms (anions) to be inserted, creating the rutile-like (TiO₂) MgH₂ structure, with a molar volume increase of 30%. Figure 2 also shows a representative example of the most stable calculated structure for the core of the larger systems.

Table 1 lists the cell parameters and atomic distances, both reference values, and the values calculated for the Mg₃₀H₆₀ cluster. The average values for the core of the cluster are close to the reference values for β -MgH₂ obtained from crystallographic data. In general, when the MgH₂ clusters contain more than 19 Mg atoms, the inner nine Mg atoms form a rutile-type

(56) Frankcombe, T. J.; Kroes, G. J.; Zuttel, A. *Chem. Phys. Lett.* **2005**, *405* (1–3), 73–78.

(57) Frish, M. J.; Pople, J. A.; Binkley, J. S. *J. Chem. Phys.* **1984**, *80* (7), 3256–3269.

(58) Becke, A. D. *J. Chem. Phys.* **1997**, *107* (20), 8554–8560.

(59) Guest, M. F.; Lenthe, J. H. v.; Kendrick, J.; Schoeffel, K.; Sherwood, P. *GAMESS-UK, A Package of Ab-Initio Programs*, version 6.0; Computing for Science Ltd.: Daresbury, UK, 2004.

(60) Cheung, S.; Deng, W. Q.; van Duin, A. C. T.; Goddard, W. A. *J. Phys. Chem. A* **2005**, *109* (5), 851–859.

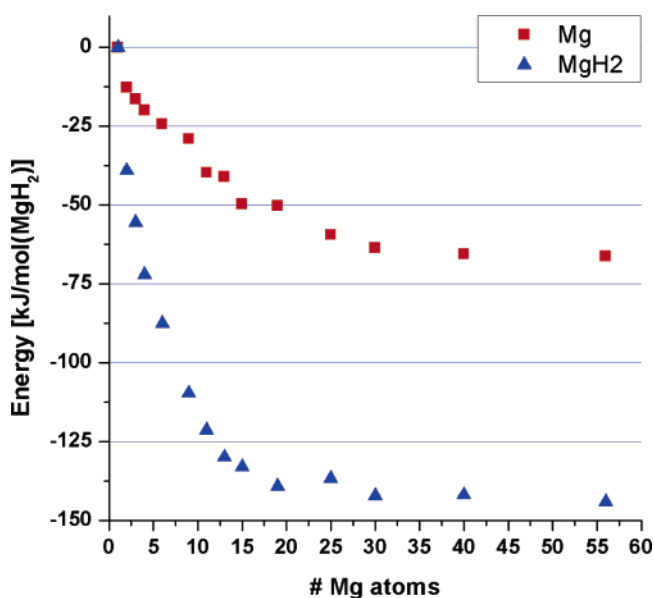
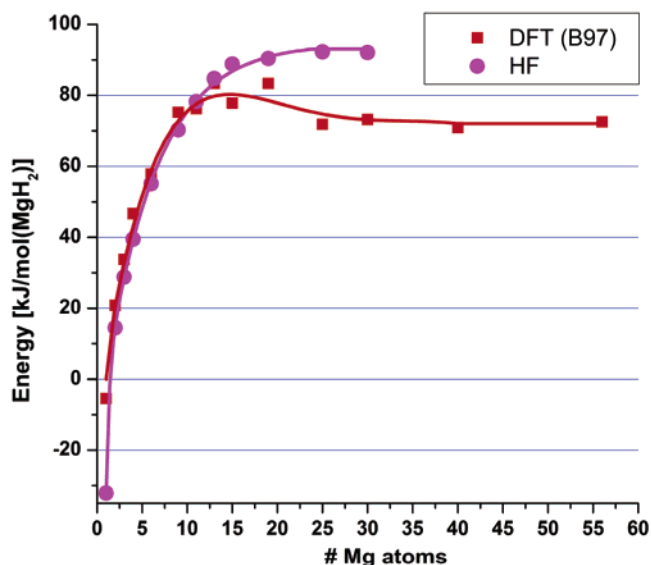
Table 1. Cell Parameters and Atomic Distances for Bulk Mg, Bulk MgH₂, and Calculated Average Values for the Core of the Mg₃₀H₆₀ Cluster^{36,37}

| cell parameters | crystal Mg | crystal β -MgH ₂ | DFT (B97) Mg ₃₀ H ₆₀ |
|--|---------------------|-----------------------------------|--|
| group | hcp, <i>p63/mmc</i> | rutile, <i>p4/mmm</i> | |
| <i>a</i> [Å] | 3.21 | 4.52 | 4.54 |
| <i>c</i> [Å] | 5.21 | 3.02 | 2.96 |
| volume per Mg(H ₂) [Å ³] | 23.2 | 30.8 | 30.5 |
| first shell Mg–H [Å] | | 1.96 | 1.97 |
| first shell Mg–Mg [Å] | 3.20 | 3.02, 3.54 | 2.96, 3.53 |
| first shell H–H [Å] | | 2.76, 3.02 | 2.74, 3.01 |

(TiO₂) structure, similar to that known for the bulk system. This result did not depend on the input geometry and was also found when starting with, for example, an hcp structure for Mg (bulk magnesium) with the hydrogen atoms placed randomly in the structure. Therefore, MgH₂ clusters with more than 19 Mg atoms can be expected to have electronic and energetic behavior similar to that of the bulk system. For Mg clusters, similar trends were reported for electronic properties, in which magnesium becomes metallic at 20 or 18 Mg atoms.^{46,49}

3.2. Energies of Mg and MgH₂ as a Function of Cluster Size. With the DFT (B97) calculations, the energies were calculated for Mg and MgH₂ clusters of up to 56 Mg atoms. These values can be found in the electronic supporting material (Supporting Information). The absolute energies were first scaled with Mg or MgH₂ as a reference energy, and subsequently divided by the number of Mg atoms in the cluster. A normalized energy per Mg(H₂) unit was thus obtained, and with this procedure the cluster energies can be conveniently compared. Figure 3 shows the calculated energies for Mg and MgH₂ as a function of cluster size.

Figure 3 shows that for both the Mg and MgH₂ clusters the DFT calculated energies depend on the number of atoms in the cluster and therefore on the cluster size. When going down in cluster size, the relative energy of the clusters becomes less negative, which indicates a destabilization of small particles. This is what can be expected from first principles of physical properties. The surface-to-volume ratio increases upon decreasing

**Figure 3.** Energies for Mg and MgH₂ clusters as a function of cluster size, calculated with the DFT method (B97 functional). The energies are scaled to the Mg or MgH₂ cluster and normalized per Mg atom.**Figure 4.** Calculated desorption energies for MgH₂ clusters with both the HF method and DFT method (B97 functional). The energies are normalized per mole of H₂ released.

ing the cluster size. Since the surface atoms have a lower coordination, the average number of bonds is lower for smaller clusters. For the metallic clusters a similar trend in desorption energy vs cluster size is found with the DFT method. However, the MgH₂ clusters are more strongly destabilized than the corresponding Mg clusters upon decreasing the cluster size below 19 Mg atoms.

3.3. Desorption Energies. With the cluster energies known, the desorption energies for MgH₂ clusters can be calculated according to eq 2. Figure 4 shows the results for both DFT and HF calculations.

The desorption energies converge to a constant value for clusters with more than 25 Mg atoms: 92.3 and 72.5 kJ·mol [H₂]⁻¹ for the HF and DFT methods, respectively. The convergence for the DFT method for clusters with more than 19 magnesium atoms is very close to the desorption enthalpy of 75 kJ·mol [H₂]⁻¹ for bulk β -MgH₂. This, and the observation that for MgH₂ clusters larger than 19 Mg atoms the core of the clusters obtains a geometry similar to that of the bulk structure, shows the validity of the DFT calculations. The deviation from the bulk value of 17 kJ·mol [H₂]⁻¹ for the HF method can be ascribed to the fact that this method does not take the correlation energy into account.

Both methods give a similar trend: the desorption energy decreases as the cluster size is reduced below 19 Mg atoms. This indicates that MgH₂ clusters with a diameter below 1.3 nm can have a desorption behavior very different from that of bulk β -MgH₂, enabling lower desorption temperatures. Upon downsizing the cluster toward the smallest stable cluster, Mg₂H₄, the desorption energy drops more than 70 kJ·mol [H₂]⁻¹. For instance, a desorption temperature of 473 K would be reached with a desorption enthalpy of 63 kJ·mol [H₂]⁻¹, corresponding to a Mg₉H₁₈ cluster size of 0.9 nm. For the smallest possible cluster, MgH₂, the desorption energy drops even to negative values, which means that the magnesium hydride molecule is not stable.

3.4. Stepwise Desorption. For clusters with 6 and 15 Mg atoms, we followed the desorption process in more detail.

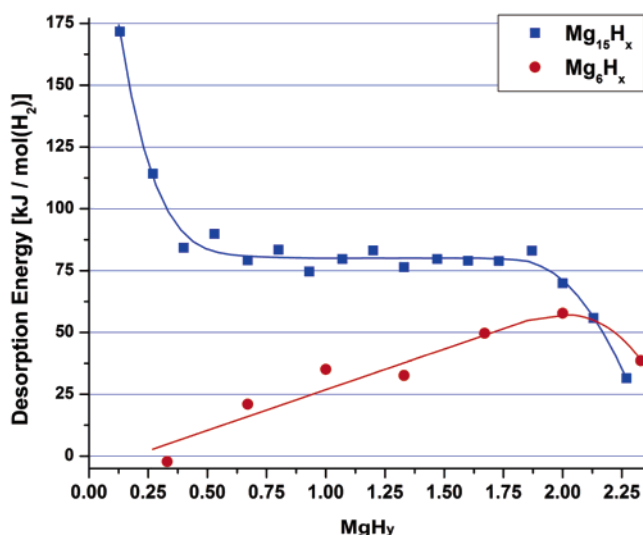


Figure 5. Energies of species involved in stepwise dehydrogenation of clusters with 6 and 15 Mg atoms, calculated with DFT (B97). The relative desorption energies are normalized per mole of H_2 released and plotted vs the H/Mg ratio. The energies for the intermediate cluster Mg_9H_x are presented in the Supporting Information.

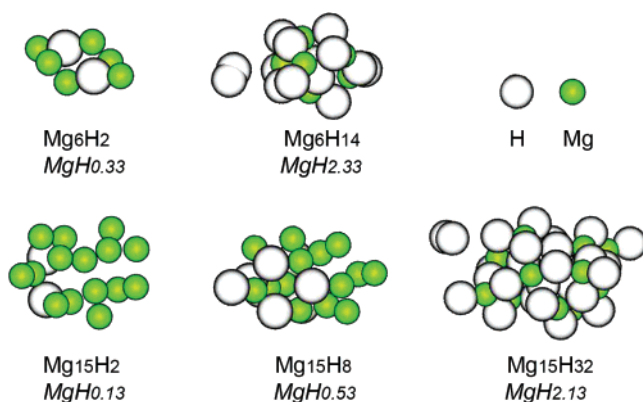
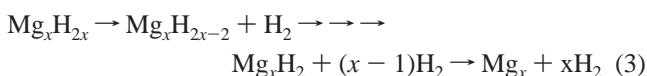


Figure 6. A few selected geometries of Mg_6H_x and $Mg_{15}H_x$ clusters.

Reaction 3 illustrates the process and the clusters involved.



For each extracted hydrogen molecule the energies of the products were calculated with the DFT (B97) method, resulting in the relative desorption energies shown in Figure 5. Figure 6 shows the corresponding geometries. Additionally, data for a Mg_9H_x cluster are given in the Supporting Information. For the $Mg_{15}H_x$ cluster, the desorption energy is constant over a broad range. From MgH_2 down to $MgH_{0.4}$ (100–20% loading) the energy needed to extract one H_2 molecule remains around a mean value of $77 \text{ kJ}\cdot\text{mol}^{-1} [H_2]^{-1}$. This value is close to experimental data for bulk MgH_2 .^{9,13,61,62} The sudden transition close to $MgH_{0.5}$ might suggest a more stable sub-hydride phase. However, a closer examination of the geometries (Figure 6, $Mg_{15}H_8$) does not support the existence of such a well-defined phase.

Going from 20% hydrogen loading downward, the calculated partial desorption enthalpy more than doubles: 174 for $MgH_{0.13}$

($Mg_{15}H_2$) versus $77 \text{ kJ}\cdot\text{mol}^{-1} [H_2]^{-1}$ for $MgH_{0.4-2.0}$ ($Mg_{15}H_{6-30}$). This indicates that the last few hydrogen molecules are more difficult to release, in this case especially the last two hydrogen molecules, which account for 13% of the hydrogen content. If we take this relatively large cluster as representative for the bulk system, this result gives a possible explanation why the theoretical 7.7 wt % reversible hydrogen uptake is usually not achieved in hydrogen absorption/desorption cycles for bulk β - MgH_2 when the applied desorption temperatures and pressures are close to the equilibrium values. Looking at the geometries in Figure 6, it is remarkable that the last two hydrogen atoms in the Mg_{15} cluster are not located in the center of the particle. In general, the hydrogen atoms tend to cluster together instead of being evenly distributed.

Hydrogen-enriched clusters were also taken into account. Two or four extra hydrogen atoms were added to the $Mg_{15}H_{30}$ cluster, leading to a potential hydrogen loading of $MgH_{2.13}$ (107%) and $MgH_{2.27}$ (113%). During the geometric optimizations these extra hydrogen atoms did not dramatically distort the $Mg_{15}H_{30}$ structure to occupy an intermetallic position, but remained on the surface. This indicates that these extra hydrogen atoms are not absorbed into, but rather adsorbed onto, the cluster, as is confirmed by the geometries of $Mg_{15}H_{32}$ and Mg_6H_{14} in Figure 6. In both cases the “excess” hydrogens are not dissociated, and are bound to the surface as a hydrogen molecule at a relatively large distance from a Mg surface atom. The possibility of this adsorption depends on the specific surface area, and is therefore less pronounced in bulk systems. Figure 5 shows that these small MgH_2 particles have the potential to take up 10–15% extra hydrogen. These extra hydrogen atoms are less strongly bound to the hydride structure and can therefore be released at lower temperatures.

For the Mg_6H_x cluster the energy trend in the stepwise desorption is distinctly different, indicating a nonbulk behavior. With a decreasing amount of hydrogen in the cluster, the desorption energy decreases. This indicates that, upon desorbing hydrogen, the Mg_6H_x cluster becomes less stable and a lower desorption temperature is needed. Also, in this case, excess hydrogen atoms can be bound to the surface as a hydrogen molecule.

3.5. Implications for Hydrogen Storage. The calculated reduction in desorption energy for clusters with fewer than 19 Mg atoms indicates a lower equilibrium desorption temperature for MgH_2 . This beneficial change in physical properties might enable the use of magnesium for hydrogen storage. The requirement of such small clusters (Mg_9H_{18} corresponds to a cluster size of 0.9 nm and a desorption temperature of 473 K) does not imply that the particles of magnesium need to be that small. Since only the crystallite size should be small, small grains could make up a larger aggregate, similar to what is observed at an order of magnitude larger scale with particles obtained with ball milling or sputtering of magnesium compounds.^{18,27,28}

Lower desorption temperatures are also reported for a special phase of magnesium hydride, γ - MgH_2 , which can be obtained by reactive ball milling of β - MgH_2 .^{9,13,29–33} For bulk systems the γ - MgH_2 is less stable than the β -form, which results in a lower desorption temperature. The calculated geometries for clusters smaller than $Mg_{13}H_{26}$ show no obvious similarity to

(61) Stampfer, J. F.; Holley, C. E.; Suttle, J. F. *J. Am. Chem. Soc.* **1960**, *82*, 3504.

(62) Zaluski, L. U.S. Patent 6,342,798, 1997.

either β - or γ -MgH₂. More importantly, because of the relative instability, the γ -structure is lost upon hydrogen cycling, being converted into β -MgH₂. In contrast, the geometries we calculated for the small clusters are the most stable geometries; hence no phase transition is expected as long as the small particle size is preserved.

As expected from the Born–Haber cycle, the downsizing of the MgH₂ and the Mg structure leads to a change in the lattice, which results in the change in desorption energy via the changed lattice energies. The corresponding possible lower desorption temperatures for clusters smaller than Mg₁₉(H₃₈) suggest a resemblance to (the structure of) amorphous Mg. It would be very interesting to experimentally verify our theoretical results for small Mg(H₂) clusters. We will report on the experimental challenge of the synthesis and stabilization of (sub)nanometer-sized Mg particles in a future paper. The possibility of making the thermodynamics of sorption more favorable by lowering the desorption energy could have a major impact on the efficiency of magnesium-based hydrogen storage materials. Furthermore, the results of our calculations can most likely be extended to other hydrogen storage materials, including complex hydrides like Mg(AlH₄)₂, and to other areas such as reactions involving nanometer-sized catalytic systems.

4. Conclusions

The energies of Mg and MgH₂ clusters of up to 56 Mg atoms were calculated with Hartree–Fock and density functional theory methods. Magnesium hydride clusters larger than Mg₁₉H₃₈ have a rutile-like geometry, similar to that of bulk β -MgH₂. The calculated desorption energy for clusters with 56 Mg atoms is roughly in agreement with the experimental bulk desorption

enthalpy of 75 kJ·mol [H₂]⁻¹.⁶¹ The calculations show that MgH₂ is more destabilized than Mg upon decreasing the cluster size below 19 Mg atoms. This translates to a decrease in desorption energy and, hence, a significantly lower hydrogen desorption temperature for these small MgH₂ clusters. A cluster of 0.9 nm would correspond to a desorption temperature of 473 K.

A stepwise calculation on the hydrogen sorption processes shows that magnesium has the potential to take up a few additional percent of hydrogen above the stoichiometric MgH₂. These extra hydrogen atoms are adsorbed to, rather than absorbed in, the hydride structure. For almost discharged bulk-like magnesium hydride clusters, the last 13% of the hydrogen is more difficult to extract. However, for the destabilized smaller clusters, hydrogen extraction becomes easier.

Our calculations clearly show that small MgH₂ clusters have a much lower desorption energy than bulk MgH₂, hence enabling hydrogen desorption at lower temperatures. At the moment, high desorption temperatures are a major impediment; hence the projected shift toward more favorable operation temperatures is crucial for the application of Mg as a reversible hydrogen storage material.

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Supporting Information Available: Desorption energies and cluster geometries. This material is available free of charge via the Internet at <http://pubs.acs.org>.

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